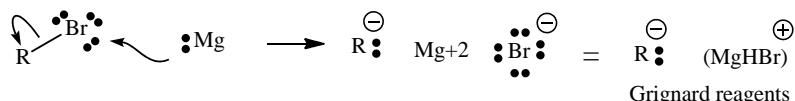


<u>Carbon electrophiles</u>	$\text{H}_3\text{C}-\text{MgBr}$ = Me-MgBr	$\text{Et}-\text{MgBr}$	$\text{i-Pr}-\text{MgBr}$	$\text{Pr}-\text{MgBr}$	$\text{c-Hx}-\text{MgBr}$	$\text{Ph}-\text{MgBr}$
$\text{H}-\text{C}(=\text{O})-\text{H}$						
$\text{Me}-\text{C}(=\text{O})-\text{H}$						
$\text{Et}-\text{C}(=\text{O})-\text{H}$						
$\text{i-Pr}-\text{C}(=\text{O})-\text{H}$						
$\text{Pr}-\text{C}(=\text{O})-\text{H}$						
$\text{c-Hx}-\text{C}(=\text{O})-\text{H}$						
$\text{Ph}-\text{C}(=\text{O})-\text{H}$						
$\text{H}-\text{C}(=\text{O})-\text{OR}$						
$\text{Me}-\text{C}(=\text{O})-\text{O}-\text{R}$						
$\text{Et}-\text{C}(=\text{O})-\text{O}-\text{R}$						
$\text{i-Pr}-\text{C}(=\text{O})-\text{O}-\text{R}$						
$\text{Pr}-\text{C}(=\text{O})-\text{O}-\text{R}$						
$\text{O}=\text{C}(=\text{O})-\text{OH}$ carboxylic acid						
$\text{O}=\text{C}(=\text{O})-\text{Cl}$						
$\text{O}=\text{C}(=\text{O})-\text{Cl}$						
$\text{H}_3\text{C}-\text{C}\equiv\text{N}$						
$\text{Et}-\text{C}\equiv\text{N}$						
$\text{i-Pr}-\text{C}\equiv\text{N}$						
$\text{Pr}-\text{C}\equiv\text{N}$						
$\text{c-Hx}-\text{C}\equiv\text{N}$						
$\text{Ph}-\text{C}\equiv\text{N}$						
O						
O						
O						

Allowed starting sources of carbon**A few key reactions in using the Grignard reaction in synthesis**

Making Grignard reagents from RBr + Mg



Making epoxides from 2 RBr (one is sulfur ylid and one is carbonyl (C=O, aldehydes or ketones)

