Textbook Errors – The page numbers listed are pages found in the solution guide

p. 7 – The first row of the transition metals in the periodic table at the top is on row too high. "Sc" (scandium) should come after "Ca" (calcium), not after "Mg" (magnesium).

p. 9 – The second row in the table of ionization energies, is a little lower than it should be. "C" should be even with and right after "B". It looks in between "B" and "Al" and could be confusing.

p. 24 – In the middle of the table in the column showing "# bonds" the third number (for oxygen) should be "2" not "4". Oxygen only makes 2 bonds when neutral. It never makes 4 bonds.

p. 29 – Problem 15, part d should show a formula of "CH_3N_3", not "CH_3NO_2", which is really part a of that problem.

p. 33 – The formula in the top, middle of the "imine" should be "CH_3CHNH". Also, there are 2 different "total electrons". The correct number should be 18 and not 20. There are 14 sigma electrons and not 16. There are still 4 extra electrons so that part is correct and the structures are correct.

p. 35, prob 17p – The correct formula should read: $NCCH(CH_3)CONH_2$ The first "N" was left out. The answer is correct in the solutions guide.

p. 38 – They syllabus lists problem 18 on p. 36, but it is really on p. 38.

p. 43 – They syllabus lists problem 19 on p. 38, but it is really on p. 42.

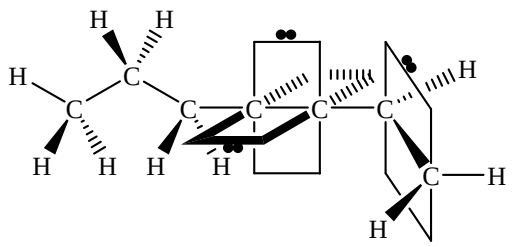
I am certain there are other errors to be found, but these are all I have written down from the textbook. Please let me know when you find others and I will post them here. Thanks.

Solution Guide Errors

p. 7, prob. 11 – carbon 2 should be trigonal planar, not tetrahedral

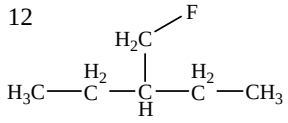
p. 10, prob 15 – d should be CH_3N_3 (CH_3NO_2 is problem 15a)

p. 15, prob 19 – forgot to draw structure l



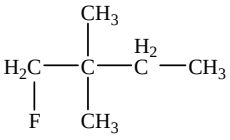
p. 29, prob 2 – started writing isomers of C7H16 instead of C6H14 on 12-18. There should only be 17 isomers, not 18. Correct structures 12-18 are shown below.

12



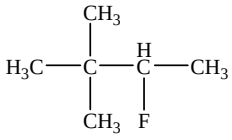
1-fluoro-2-ethylbutane

13



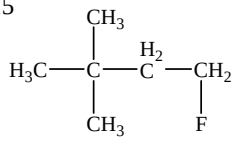
1-fluoro-2,2-dimethylbutane

14



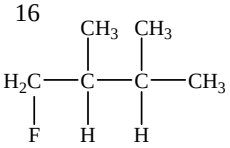
2-fluoro-3,3-dimethylbutane

15



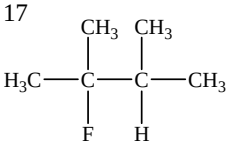
1-fluoro-3,3-dimethylbutane

16



1-fluoro-2,3-dimethylbutane

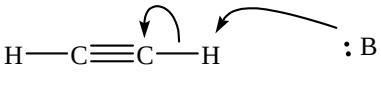
17



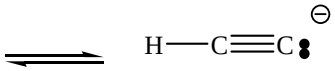
2-fluoro-2,3-dimethylbutane

p. 73, prob 4b – the conjugate base is drawn incorrectly. It should look as below.

b.



$K_a = 10^{-25}$

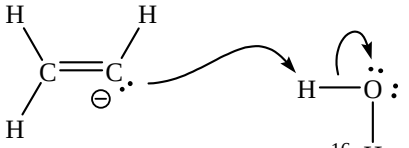


$$\frac{K_a\,(HCCH) = 10^{-25}}{K_a\,(H_2O) = 10^{-16}} = 10^{-9}$$

Weaker than water.

p. 75, prob 7c – the conjugate acid is drawn incorrectly. It should look as below.

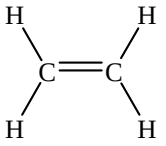
c.



Stronger base.

$K_a = 10^{-16}$

$$K_{eq} = \frac{K_a = 10^{-16}}{K_a = 10^{-44}} = 10^{28}$$



$K_a = 10^{-44}$