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Calculating ¹H NMR Chemical Shifts

If a carbon has more than one functional group directly attached to it, the following table can be used to estimate the chemical shift of an attached hydrogen.

$$\text{—CH}_2\text{—} \quad \text{Calculation} \quad \delta \text{ (ppm)} \quad \text{R}_1\text{—CH}_2\text{—R}_2 = 1.2 + \Delta_{\text{R}_1} + \Delta_{\text{R}_2}$$

Methylenes (CH₂) have two groups attached, so the starting chemical shift (1.2 ppm) will be adjusted using two values from the table (Δ values).

$$\begin{array}{c} | \\ \text{—CH} \\ | \end{array} \quad \text{Calculation} \quad \delta \text{ (ppm)} \quad \begin{array}{c} \text{R}_1 \\ | \\ \text{R}_2\text{—}\dot{\text{C}}\text{—H} \\ | \\ \text{R}_3 \end{array} = 1.5 + \Delta_{\text{R}_1} + \Delta_{\text{R}_2} + \Delta_{\text{R}_3}$$

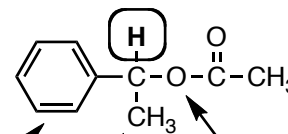
Methines (CH) have three groups attached, so the starting chemical shift (1.5 ppm) will be adjusted using three values from the table (Δ values).

Protons on sp³ Carbons: Chemical Shift Calculation Table

—R	Δ	Description	—R	Δ	Description
—CH ₂ R	0.0	alkyl	—OH	2.3	α to alcohol O
—CR=CR ₂	0.8	allylic (next to alkene)	—OR	2.1	α to ether O
—C≡CR	0.9	propargylic (next to alkyne)	—OAr	2.8	α to O of aromatic ether
—C≡N	1.2	α to cyano	$\text{—O—}\overset{\text{O}}{\parallel}{\text{C}}\text{—R}$	2.8	α to O of ester
—Ar	1.4	benzylic	$\text{—O—}\overset{\text{O}}{\parallel}{\text{C}}\text{—Ar}$	3.1	α to O of aromatic ester
$\text{—}\overset{\text{O}}{\parallel}{\text{C}}\text{—R}$	1.2	α to ketone or aldehyde C=O	—NR ₂	1.5	α to amine
$\text{—}\overset{\text{O}}{\parallel}{\text{C}}\text{—OR}$	1.1	α to ester or carb. acid C=O	—NO ₂	3.2	α to nitro
$\text{—}\overset{\text{O}}{\parallel}{\text{C}}\text{—Ar}$	1.7	α to C=O of aromatic ketone	—SR	1.3	α to thiol or thioether
—F	3.2	α to fluorine			
—Cl	2.2	α to fluorine			
—Br	2.1	α to bromine			
—I	2.0	α to iodine			

Example:

estimate the chemical shift of the selected proton



$$\begin{aligned} \delta &= 1.5 + \Delta(-\text{Ar}) + \Delta(-\text{R}) + \Delta(-\text{OCOR}) \\ \delta &= 1.5 + 1.4 + 0.0 + 2.8 \\ \delta &= 5.7 \text{ ppm (actual 5.4 ppm)} \end{aligned}$$